

THE AIChE CONCEPT WAREHOUSE: A WEBSITE RESOURCE FOR CONCEPT-BASED INSTRUCTION

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Many engineering educators and industry partners emphasize the need for students to apply their knowledge to new and challenging problems. A lack of conceptual understanding has been shown to severely restrict students' ability to solve new problems since they do not have the foundational understanding to use their knowledge in new situations.^[1] This problem is exacerbated by science and engineering classrooms that often reward students more for rote learning than for conceptual understanding.^[2] There is clearly a need for more emphasis on conceptual understanding in the chemical engineering curriculum. But developing high-quality instructional materials is time-consuming and requires expertise.^[3] Even with access to instructional materials, it is also useful to be able to deliver them in class in ways that all students can participate and that both students and the instructor get formative feedback.

The AIChE Concept Warehouse (CW) is a community-developed, web-based tool to decrease instructional barriers and to help faculty implement concept-based active learning in class.^[4,5] The CW provides three distinct but complementary functions: (a) a content repository, (b) an audience response system and learning management system to deliver content, and (c) learning analytics that provide learning data to instructors and researchers. Instructor and student interfaces are available free at <<http://cw.edudiv.org>>. After obtaining a CW account, instructors can navigate through the Instructor

Interface to find content, deliver it in class through the Student Interface, and use reported aggregate and individual answers and scores to guide instruction. A video describing the CW is available at <<https://youtu.be/Nf5w0kG3asY>>.

I next briefly describe the three major functions of the CW.

CONTENT REPOSITORY

Educators have access to content through the Instructor Interface, which is organized by tabs. There are three content tabs available: Conceptests, Concept Inventories, and Instructional Tools:

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Conceptests: Figure 1 shows two screenshots of the “Conceptests” tab. Conceptests are conceptual questions that ask students to use a core concept or set of concepts to reason through to the correct answer. They help students develop understanding of engineering concepts and help faculty and students identify the level of mastery.

As shown in Figure 1 (a), instructors currently have access to around 3,000 Conceptests available for core chemical engineering classes (see list on left of figure). These can be searched by class, by topic, or using the advanced search tool. Instructors can also write their own Conceptests [Figure 1 (b)], which they can choose to either contribute to the community (for a well-written and general question) or reserve for use only in their own classes (for a question specific to the instructor’s delivery). There are several question types to select from including multiple-choice with single answer, multiple-choice with multiple answers, short answer, and ranking exercises.

There are two ways instructors can use this content. Figure 2 shows a screenshot of the Student Interface of a Conceptest on the CW website. This example shows a multiple-choice question with a figure. When assigning a Conceptest, the instructor has the option to choose to have students provide a written explanation and/or answer how confident they are when answering the question (both selected here). Asking students to provide written explanations has been shown to increase student learning.^[6-8] To deliver content through the CW, the instructor must set up a class using the Classes tab (see section below). Alternatively, the questions can be downloaded from the CW in PowerPoint or Word for other uses.

Concept Inventories: There are 12 concept inventories (CIs) available for summative assessment of students’ conceptual understanding. CIs are test instruments meant to help instructors and researchers determine the extent of their students’ conceptual understanding about a specific subject.^[9,10] The CIs available on the CW have been developed and psychometrically tested according to validity and reliability criteria. These instruments focus on a few core concepts in a subject and several questions (items) are used to test understanding of each concept. CIs can be used in ABET assessment processes or in rigorous education research studies by administering the CI at the start and end of a term to quantify learning gains.

Figure 3 shows a screenshot of the Concept Inventory

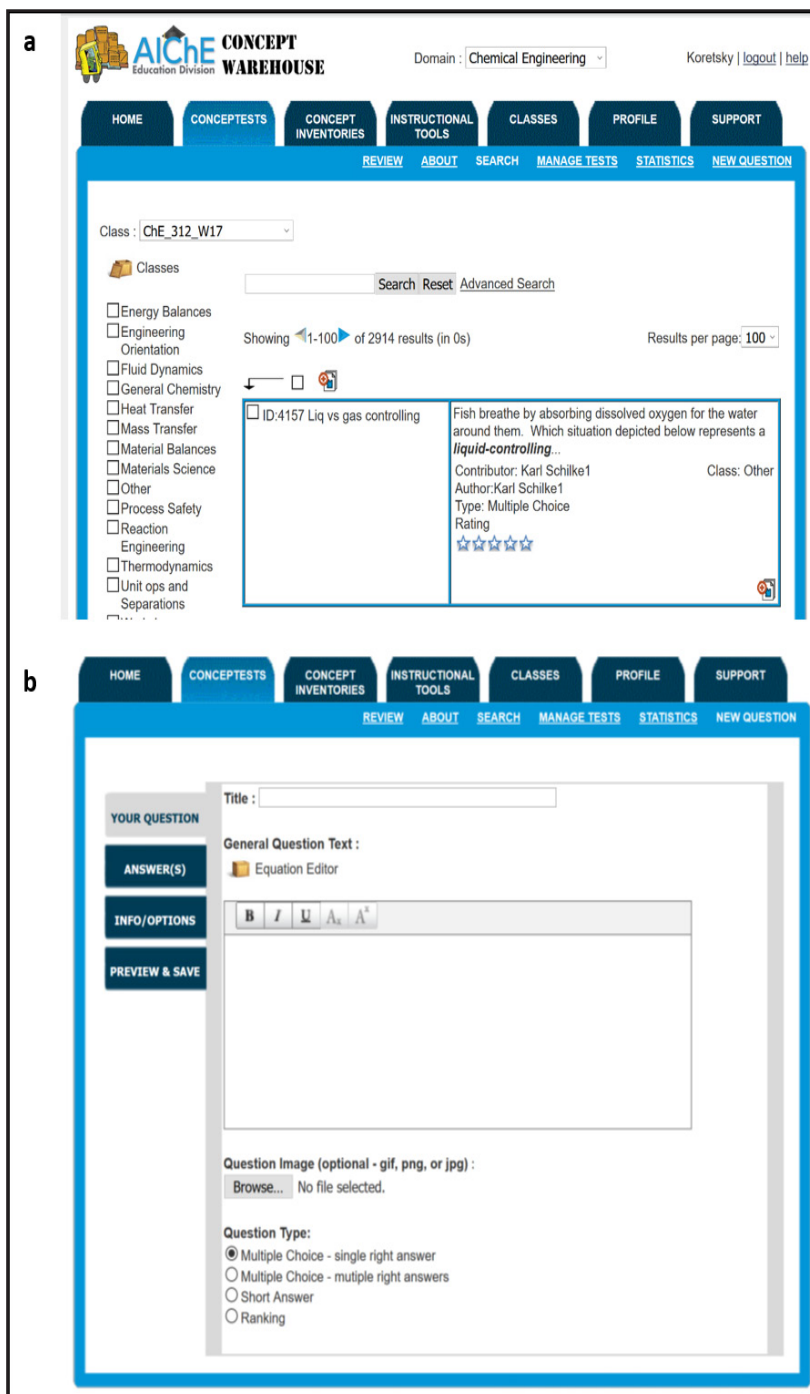


Figure 1. Instructor Interface: Screenshots of the Conceptests “tab” in the Instructor Interface: (a) the search menu; (b) the new questions tab.

tab for the Heat and Energy Concept Inventory (HECI).^[11] As shown, an instructor can see the list of items in the “Questions” menu (on left) or select the “Answer Key” to get the same list, but with the answer highlighted. In this way, the instructor has the opportunity to work through the test instrument before assigning it. The following CIs

HOME
QUESTIONS
PROFILE

Question text

Plus Figure

Written Reflection (optional)

Confidence (optional)

An ideal gas flows steadily through the piping system and valve shown below. The inlet pressure and temperature are P_1 and T_1 and the pressure drops through the valve to a lower value, P_2 .

Assuming the valve is well insulated and inlet and outlet pipes connected to the valve are the same diameter, what is the relationship of the outlet temperature T_2 to the inlet temperature T_1 ?

$T_2 > T_1$ because work is done on the gas as it is compressed through the valve opening
 $T_2 < T_1$ because temperature must decrease if pressure decreases since the volume and number of moles both stay the same
 $T_2 = T_1$ because rapid expansion of an ideal gas does not affect temperature
 Can't answer unless the type of gas flowing is specified

Please explain your answer in the box below.

Please rate how confident you are with your answer.

substantially
unsure

moderately
unsure

neutral

moderately
confident

substantially
confident

Submit

Multiple
Choice
Answer
Selections

Figure 2. Student Interface: an example of a Conceptest delivered to students.

are available.

- *Thermodynamics* (35 items)
- *Heat and Energy* (36 items)^[11]
- *three Thermal and Transport Concept Inventories (TTCI)*^[12].
TTCI: *Thermodynamics* (24 items); TTCI: *Fluids* (26 items); and TTCI: *Heat Transfer* (18 items)
- *Materials* (30 items)^[13]
- *Statistics* (25 items)^[9]
- *Dynamics* (29 items)^[14]
- *four Chemistry CIs* (20-31 items)^[15]

Instructional Tools: In addition to Conceptests and CIs, we have recently been adding a set of Instructional Tools to help students develop conceptual understanding. Three of these tools are described next.

Reflection Activities: Quick end-of-class activities (often called minute papers or exit slips), where students are asked to reflect on class materials, have been described as providing one of the best returns on investment of class time and instructor effort.^[16] These activities provide a way

to help students reflect on content and provide instructors an assessment of where students are struggling. Based on research findings,^[17] the CW provides a hybrid reflection activity that allows students to choose amongst a “Most Surprised” prompt, “Muddiest Point” prompt, or both. The Instructor Interface provides a list of student responses, which in a large class can be overwhelming. To facilitate interpretation, the instructor is also provided word clouds to analyze the data, and a table of the most common response themes.

Interactive Virtual Labs: In the Interactive Virtual Laboratories (IVLs) students are guided through a sequence of around 15 “frames” that prompt them to interact with simulations and relate microscopic (molecular) phenomena of a system to macroscopic representations and answer questions.^[18] The IVLs target difficult concepts, and each IVL draws upon known student difficulties from the science and engineering education literature. Figure 4 shows the Heat Capacity IVL, which helps students understand why constant pressure heat capacity (c_p) is larger than constant volume heat capacity (c_v). The CW has the following IVLs available: PV Work; Reversibility; Heat Capacity (c_v vs. c_p); Hypothetical Paths/Chemical Reaction;

Figure 3. Instructor Interface: a screenshot of the concept inventory tab.

Hypothetical Paths/Phase Equilibrium; Reaction Rate and Chemical Equilibrium; and Phase Equilibrium. A video describing the IVLs is available at <https://youtu.be/1t17fh026Pw>.

Inquiry-Based Activities: Inquiry-based activities (IBAs) have been shown to significantly improve student scores on tests of conceptual learning.^[19,20] The CW currently has two heat transfer IBAs on topics in which students commonly show misconceptions: rate vs. amount and radiation. As shown in Figure 5, the Rate vs. Amount IBA helps students understand the difference between how fast energy is transferred and the total amount of energy transferred. The recommendations in Figure 5 on effectiveness and the perceived ease of use are based on research data.

SITE FOR INTERACTIVE TEACHING (AUDIENCE RESPONSE SYSTEM)

As an instructor, you can use the site in class or for homework to deliver content and collect responses, as students answer on their devices (cell phones, tablets, laptops). Figure 6 shows screenshots of the mobile student app, which is formatted to optimize use on cell phones. Setting up the CW to use in class is a one-time process (for each class you teach) and requires only a few steps: creating a new class in the “Classes” tab, pasting in the emails of your students, and sending

Figure 4. Heat capacity (c_v vs. c_p) Interactive Virtual Laboratory. The actual IVL image has been rendered with a white background for readability.

an automated email that provides them a link to the Student Interface. When you find a question, inventory, or tool that you want to use, add it to “Manage Tests,” “Manage Inventories,” or “Manage Tools”; and then assign it from the menu. You can also link several Conceptests together in “Manage Tests” so that they are delivered sequentially in a single assignment.

ANALYTICS FOR FORMATIVE ASSESSMENT

If you assign activities through the CW, you can take advantage of the analytic and scoring capability on the Instructor Interface. Figure 7 shows screenshots of analytics available including a catalog of student responses throughout the term (left) and multiple-choice and written responses to a specific conceptest (right). When you ask students to provide written explanations or you use the Surprised and Muddy reflections Instructional Tool, clickable word clouds are available to help make sense of student responses more quickly.^[21] You can also access a grade sheet that summarizes student performance.

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The screenshot shows the AIChE Concept Warehouse interface. The main content area displays the 'Heat Transfer - Rate vs. Amount IBAs' activity. Below the activity description, there is a list of delivery options for 'Activity 1 (Cooling Beverages) Delivery Options'. A table provides details for five options, including their type, delivery mode, effectiveness, and ease of use.

Option	Type	Delivery Mode	Effectiveness (Explanation)	Ease of Use (Explanation)
1	Physical Experiment	Performed by students	High (details)	Higher Effort (details)
2		Instructor demonstration	Medium (details)	Moderate Effort (details)
3	Simulation	Performed by students	Medium (details)	Low Effort (details)
4		Instructor demonstration	Medium (details)	Low Effort (details)
5	Thought Experiment	Instructor-led discussion	Low (details)	Low Effort (details)

Figure 5. The Rate vs. Amount Inquiry-Based Activity. There are five delivery options available (bottom right).

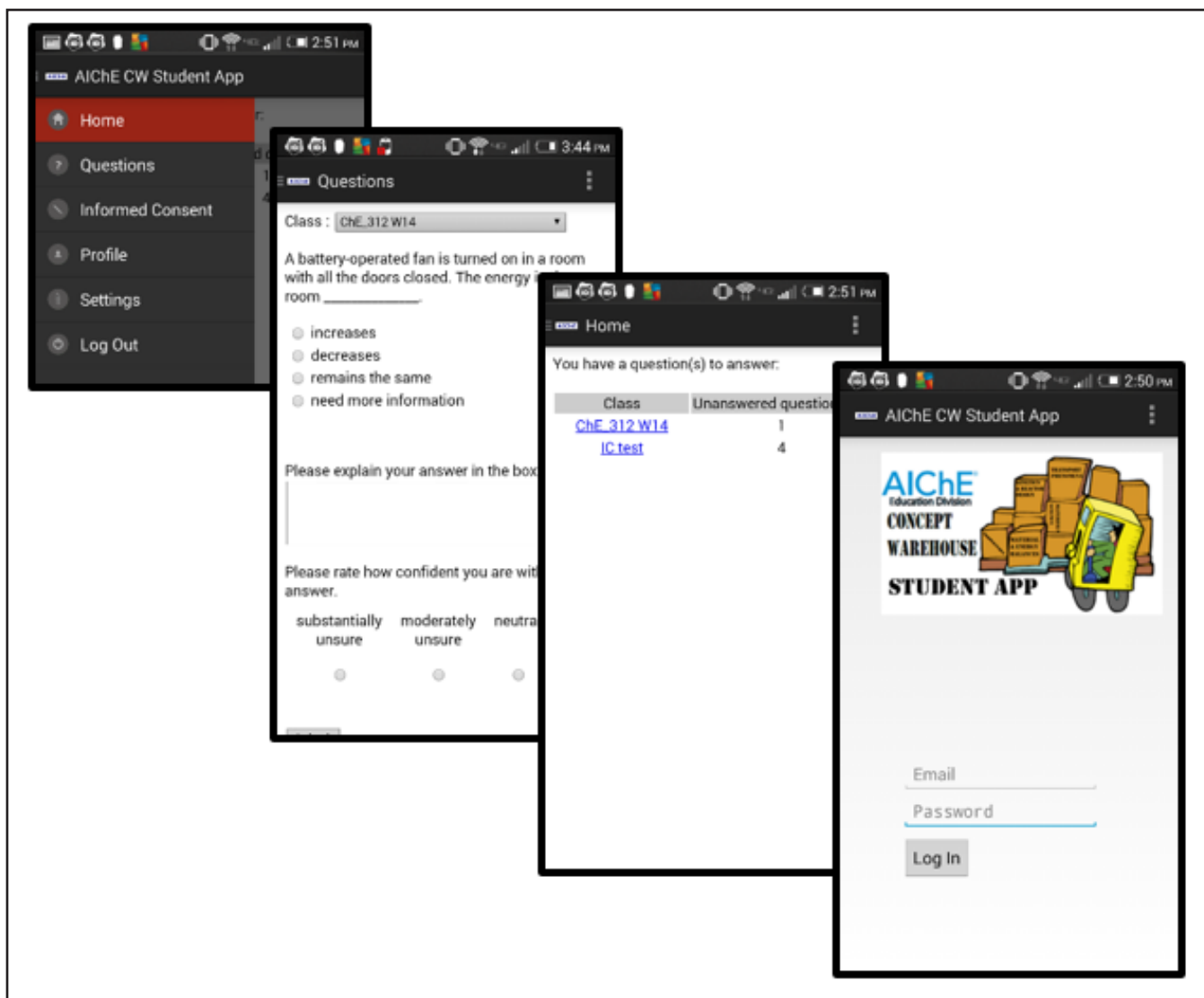


Figure 6. Screenshots of the Student Interface on mobile student app for interactive use in class.

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HOME CONCEPT TESTS CONCEPT INVENTORIES INSTRUCTIONAL TOOLS CLASSES PROFILE

ABOUT SEARCH MANAGE TESTS STATISTICS

Class: Thermodynamics

Show Page Options

1187 Pure Species Phase Equilibrium - T	95.8 %	92-2.2	N/A	96 Ans/96 Stu
1188 Pure Species Phase Equilibrium - Ca	75.8 %	72-22.1	N/A	95 Ans/96 Stu
Original Throttling Valve	20.6 %	20-70.7	Neutral	97 Ans/97 Stu
Ideal gas - enthalpy	62.2 %	61-29.8	Neutral	98 Ans/97 Stu
Ideal gas - enthalpy	41.7 %	40-40.16	Neutral	96 Ans/97 Stu

Original Throttling Valve

An ideal gas flows steadily through the piping system and valve shown below. The inlet pressure and temperature are P_1 and T_1 and the pressure drops through the valve to a lower value, P_2 .

Assuming the valve is well insulated and inlet and outlet pipes connected to the valve are the same diameter, what is the relationship of the outlet temperature T_2 to the inlet temperature T_1 ?

$T_2 > T_1$ because work is done on the gas as it is compressed through the valve opening

$T_2 < T_1$ because temperature must decrease if pressure decreases since the volume and number of moles both stay the same

$T_2 = T_1$ because rapid expansion of an ideal gas does not affect temperature

Can't answer unless the type of gas flowing is specified

Figure 7. Screenshots of analytics available on the Instructor Interface.

Short Answer Explanation Word Cloud

ideal gas pressure decrease temperature volume constant drop energy law system mass

Short Answer Explanation Word Cloud (by answer choice)

$T_2 < T_1$ because temperature must decrease if pressure decreases since the volume and number of moles both stay the same	ideal gas pressure decrease temperature	PV system number moles volume constant law drop stay change work
$T_2 = T_1$ because rapid expansion of an ideal gas does not affect temperature	ideal gas energy temperature	system constant change pressure energy work
$T_2 > T_1$ because work is done on the gas as it is compressed through the valve opening	inlet ideal gas temperature pressure	

Answer(s)	Explanation	Confidence
$T_2 = T_1$ because rapid expansion of an ideal gas does not affect temperature	It is an expanding gas, and expanding gases get cooler. T but volume	2
$T_2 = T_1$ because temperature must decrease if pressure decreases since the volume and number of moles both stay the same	Since the outlet volume is the same and the ideal gas law gives a relation of $PV = RT$, as P decreases T must also decrease.	3
$T_2 = T_1$ because temperature must decrease if pressure decreases since the volume and number of moles both stay the same	With a valve there is no work being done on the system eliminating 1. There is a change in the intrinsic variables so there must be a response in the other variables to balance out the system eliminating 3. If there is a pressure drop the molecules will become less packed reducing kinetic energy and removing heat in the process.	4
$T_2 = T_1$ because temperature must decrease if pressure decreases since the volume and number of moles both stay the same		5

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